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Thorax location human body

As people get closer to the time that astronauts and explorers will live and work in space for extended periods of time, many questions arise about what it will be like for those who have their careers out there. There's a lot of data based on long-haul flights from astronauts like Mark Kelly and Peggy Whitman, but the life sciences experts of most space agencies need a lot more data to understand what will happen to future travelers. They already know that the long-term inhabitants aboard the International Space Station have experienced a number of important and puzzling changes in their bodies, some of which last long after they return to Earth. Mission planners use their experiences to plan missions to the moon, Mars and beyond. NASA However, despite this priceless data from actual experiences, people also get a lot of non-valuable data from Hollywood movies about what it's like to live in space. In those cases, drama usually trumps scientific accuracy. In particular, the movies are great on gore, especially when it comes to portraying the experience of being exposed to vacuum. Unfortunately, those movies and TV shows (and video games) give the wrong impression about what it's like to be in space. In the 1981 film Outland, starring Sean Connery, there is a scene in which a construction worker in space gets a hole in his suit. As the air leaks out, the internal pressure drops and his body is exposed to a vacuum, we look in horror through his front plate as it swells and explodes. Could that really happen, or was that dramatic license? A somewhat similar scene occurs in the 1990 Arnold Schwarzenegger film, Total Recall. In that film, Schwarzenegger abandons the pressure of the habitat of a Mars colony and begins to inflate like a balloon in the much lower pressure of the Martian atmosphere, not quite a vacuum. He is saved by creating a whole new atmosphere by an ancient alien machine. Again, can that happen, or was dramatic license in the game? These scenes raise a perfectly understandable question: What happens to the human body in a vacuum? The answer is simple: it won't explode. The blood doesn't boil either. However, it will be a quick way to die if an astronaut's spacesuit is damaged. There are a number of things about being in space, in a vacuum, that can harm the human body. The unfortunate space traveler would not be able to hold their breath for long (if at all) because it would cause lung damage. The person would likely remain conscious for several seconds until the blood without oxygen reaches the brain. Then all bets are off. The vacuum of space is also quite darn cold, but the human body does not lose heat that fast, so Unfortunate astronaut would have a little time before freezing to death. It's possible they might have some problems with their eardrums, including a fracture, but maybe not. Stranded in space exposes the astronaut to high radiation and the chance of a bad sunburn. Their bodies might actually swell some, but fail to recall the proportions so dramatically displayed in Total. The bends are also possible, just like what happens to a diver who shows up too fast from a deep underwater dive. That condition is also known as decompression disease and happens when dissolved gases in the bloodstream create bubbles as the person decompresses. The condition can be fatal and is taken seriously by divers, high-altitude pilots and astronauts. NASA/Bill Stafford/Wikimedia Commons/Public Domain Although normal blood pressure will make a person's blood boil, the saliva in their mouth might very well start to do so. There's evidence of that happening from an astronaut who's been through it. In 1965, while conducting tests at the Johnson Space Center, a subject was accidentally exposed to a near vacuum (less than a psi) when its spacesuit leaked while in a vacuum chamber. He did not faint for about fourteen seconds, by which time unoxygenated blood had reached his brain. Technicians began suppressing the room within fifteen seconds and he regained consciousness around the equivalent of 15,000 feet altitude. He later said that his last conscious memory was of the water on his tongue starting to boil. So there is at least one data point about what it's like to be in a vacuum. It won't be pleasant, but it won't be like the movies either. There have been cases of parts of astronauts' bodies being exposed to vacuum when suits were damaged. They survived due to quick action and safety protocols. The good news of all those experiences is that the human body is incredibly resilient. The worst problem would be lack of oxygen, not lack of pressure in the vacuum. If returned to a normal atmosphere fairly quickly, a person would survive with few or no irreversible injuries after accidental vacuum exposure. More recently, astronauts on the International Space Station found an air leak from a hole created by a technician on the ground in Russia. They weren't in danger of losing their air right away, but they had to go to some trouble to keep it safe and permanently connected. Edited and updated by Carolyn Collins Petersen. There are several ways to consider the composition of the human body, including the elements, the type of molecule, or the type of cells. Most of the human body consists of water, H2O, with bone cells consisting of 31% water and the lungs 83%. Carbon, the basic unit for organic molecules, comes second. 96.2% of the mass of the human body consists of only four elements: oxygen, carbon, hydrogen and nitrogen. Oxygen (O) - 65% - Oxygen along with hydrogen form water, which is the primary solvent found in the body is used to regulate temperature and osmotic pressure. Oxygen is found in many important organic compounds. Carbon (C) - 18.5% - Carbon has four adhesion sites for other that makes it the most important atom for organic chemistry. Carbon chains are used to build carbohydrates, fats, nucleic acids and proteins. Breaking bonds with carbon is an energy source. Hydrogen (H) - 9.5% - Hydrogen is found in water and in all organic molecules. Nitrogen (N) - 3.2% - Nitrogen is found in proteins and in the nucleic acids that are part of the genetic code. Calcium (Ca) - 1.5% - Calcium is the most common mineral in the body. It is used as a structural material in bones, but it is essential for protein regulation and muscle contraction. Phosphorus (P) - 1.0% - Phosphorus is found in the molecule ATP, the primary energy carrier in cells. It's in the bone, too. Potassium (K) - 0.4% - Potassium is an important electrolyte. It is used to transfer nerve impulses and heart rate regulation. Sodium (Na) - 0.2% - Sodium is an important electrolyte. Like potassium, it is used for nerve signaling. Sodium is one of the electrolytes that helps regulate the amount of water in the body. Chlorine (Cl) - 0.2% - Chlorine is an important negatively charged ion (anion) that is used to maintain moisture balance. Magnesium (Mg) - 0.1% - Magnesium is involved in more than 300 metabolic reactions. It is used to build the structure of muscles and bones and is an important cofactor in enzymatic reactions. Sulfur (S) - 0.04% - Two amino acids are sulfur. The bonds sulfur forms help shape proteins to the shape they need to perform their functions. Many other elements can be found in very small quantities (less than 0.01%). For example, the human body often contains traces of thorium, uranium, samarium, tungsten, beryllium and radium. Trace elements considered essential in humans include zinc, selenium, nickel, chromium, manganese, cobalt and lead. Not all elements found in the body are essential for life. Some are considered contaminants that seem to do no harm, but serve no known function. Examples include caesium and titanium. Others are actively toxic, including mercury, cadmium and radioactive elements. Arsenic is considered toxic to humans, but serves a function in other mammals (goats, rats, hamsters) in traces. Aluminum is interesting because it is the third most common element in the Earth's crust, but its role in the human body is unknown. While fluorine is used by plants to produce protective toxins and has apparent beneficial intake in humans. You also view the elementary composition of an average human body per mass. Chang, Raymond (2007). Chemistry, 9th edition. Mr. McGraw-Hill. ISBN 0-07-110595-6.Emsley, John (2011). The building blocks of nature: an A-Z guide to the elements. OUP Oxford. p. 83. ISBN 978-0-19-960563-7.Frausto Da Silva, J.J.R. Williams, R.J. P (2001-08-16). The chemie van de elementen: de anorganische chemie van het leven. ISBN 9780198508489.H. A., V. W. Rodwell; P. A. 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